DR. RAM MANOHAR LOHIA AVADH UNIVERSITY, AYODHYA

Structure of Syllabus for the

Program: M.Sc., Subject: CHEMISTRY

Structure of Syllabus Developed by				
S.N.	**Name of BoS Convener/ BoS Member**	**Designation**	**Department**	**College/ University**
1.	Prof. D. N. Singh	Professor & Convener	Department of Chemistry	K. S. Saket PG College Ayodhya, U.P.
2.	Dr. Ram Vishun Prasad	Assistant Professor & Internal Expert	Department of Chemistry	A.N. D. Kishan PG College Babhnan, Gonda, U.P.
3.	Prof. Nanhai Singh	Ex-Professor & External Expert	Institute of Science, Department of Chemistry	BHU Varanasi, U.P.
4.	Prof. Lal Bahadur	Ex-Professor & External Expert	Institute of Science, Department of Chemistry	BHU Varanasi, U.P.
5.	Prof. D.S. Pandey	Professor & External Expert	Institute of Science, Department of Chemistry	BHU Varanasi, U.P.
6.	Prof. Abha Bishnoi	Professor & External Expert	Department of Chemistry	University of Lucknow, U.P.
7.	Prof. S.S. Yadav	Ex-Professor & External Expert	Department of Chemistry	DDU University Gorakhpur, U.P.

5. Prof. C.K. Mishra Dean Science

Dr. RML Avadh University

03.09.22

03/09/2022

Program: M.Sc., Subject: CHEMISTRY

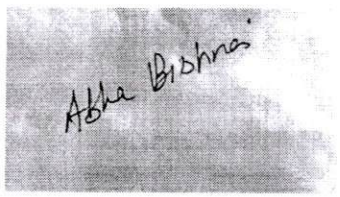
Course Code		Course Title	Credits	T/P	Evaluation	
					CIE	ETE
A	B	C	D	E	F	G
SEMESTER I (YEAR I)						
B020701T	CORE	Spectroscopy-I	5	T	25	75
B020702T	CORE	Symmetry & Molecular Vibrations	5	T	25	75
B020703T	CORE	Organic Reaction Mechanism	5	T	25	75
B020704T	FIRST ELECTIVE (Select any one)	Chemical Kinetics & Thermodynamics	5	T	25	75
B020705T		Surface & Solid-State Chemistry	5	T	25	75
B020706P	SECOND ELECTIVE (Select any one)	Chemistry Laboratory Course-I A	5	P	50	50
B020707P		Chemistry Laboratory Course-I B	5	P	50	50
SEMESTER II (YEAR I)						
B020801T	CORE	Chemistry of Main Group Elements	5	T	25	75
B020802T	CORE	Stereochemistry & Spectroscopy-II	5	T	25	75
B020803T	CORE	Advanced Quantum Mechanics	5	T	25	75
B020804T	THIRD ELECTIVE (Select any one)	Research Aptitude	5	T	50	50
B020805T		Environmental Science	5	T	50	50
B020806P	FOURTH ELECTIVE (Select any one)	Chemistry Laboratory Course-II A	5	P	50	50
B020807P		Chemistry Laboratory Course-II B	5	P	50	50
SEMESTER III (YEAR II)						
B020901T	CORE	Coordination & Bioinorganic Chemistry	5	T	25	75
B020902T	CORE	Pericyclic, Photochemistry & Rearrangement Reactions	5	T	25	75
B020903T	CORE	Electrochemistry	5	T	25	75
B020904T	FIFTH ELECTIVE (Select any one)	Natural Products	5	T	25	75
B020905T		Medicinal Chemistry	5	T	25	75
B020906P	SIXTH ELECTIVE (Select any one)	Chemistry Laboratory Course-III A	5	P	50	50
B020907P		Chemistry Laboratory Course-III B	5	P	50	50

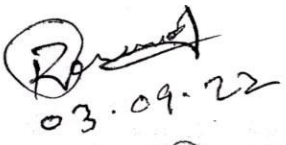
SEMESTER IV (YEAR II)						
B021001T	CORE	Organotransition Metal Chemistry	5	T	25	75
B021002T	CORE	Organic Synthesis	5	T	25	75
B021003T	SEVENTH ELECTIVE (Select any one)	Analytical Chemistry	5	T	25	75
B021004T		Polymer Chemistry	5	T	25	75
B021001P	RESEARCH PROJECT/ DISSERTATION	Practical Based Major Research Project /Dissertation	10	P	50	50

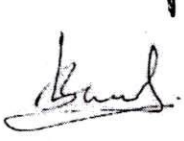
NOTE:

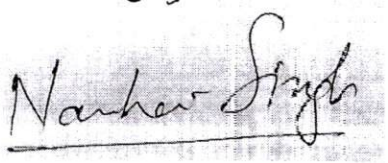
1. **T/P** in Column-E stands for **Theory/Practical.**
2. **CIE** in Column-F stands for **'Continuous Internal Evaluation'** and depicts the maximum internal marks. Respective examination will be conducted by subject teacher.
3. **ETE** in Column-G stands for **'External Evaluation'** and depicts the maximum external marks. Respective Examination will be conducted by the University.
4. Column-B defines the nature of course/paper. The word **CORE** herein stands for **Compulsory Subject Paper**.
5. Column-D depicts the credits assigned for the corresponding course/paper.
6. **First Elective:** It will be a Subject Elective. Students may select one of the two subject papers under this category.
7. **Second Elective:** It will designate a Practical Paper or equivalently a Field Visit or Project Presentation. In case of Field Visit, student is required to submit a detailed report of the visit for the purpose of evaluation. The report should include the observational features and benefits of the visit. In case of Project Presentation, the student may be assigned to go for a survey/practical or theoretical project/assignment or seminar with presentation.
8. **Third Elective:** It will be a Generic Elective. The student may study or receive training of the any subject of his interest (depends on the availability in his institution of enrollment). The Generic elective paper will be evaluated in two parts, first part (50 marks) would be a continuous internal evaluation (03 tests 20+20+10 marks) whereas the examination and evaluation of the second part (50 marks) would be arranged by the college itself (01 exam).
9. **Fourth Elective:** It will accommodate a practical paper or Industrial Training or Project Presentation. In case of Industrial Training, student may be allowed for the summer training and is required to submit a detailed training report including training certificate for the evaluation.
10. **Fifth Elective:** It will be a Subject Elective. Students may select one of the two subject papers under this category.

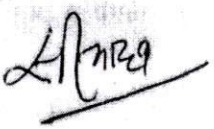
11. **Sixth Elective:** It will be a Practical Paper or equivalently a Project Presentation based on Survey/ Seminar/ Assignment. In case of Project Presentation, student has to submit an exhaustive report on respective topic and to face an open presentation for the evaluation.
12. **Seventh Elective:** It will be a Major Research Project or equivalently a research-oriented Dissertation on the allotted topic. The student straight away will be awarded 05 credits if he publishes a research paper on the topic of Research Project or Dissertation.
13. There will be a Major Research Project or equivalently a research oriented Dissertation on Semester-IV. The Student straight away will be awarded 05 Credits if he publishes a research paper on the topic of the research Projector Dissertation.
14. Methodology for the practical examination and examiner appointment will be governed by the Clause-13 of the NEP Guideline of RMLAU dated 27-06-2022 except the marks distribution for continuous internal evaluation and external evaluation.











Subject Prerequisites:
To study this course, a student must have passed B.Sc. III Year, VI Semester with Chemistry as a major subject.

Program Outcomes (POs):
• The program has been designed to enable the students to acquire strong theoretical, practical and research knowledge in the various areas of Chemistry. • The program covers maximum branches of Chemistry and Experimental Laboratory Courses as well as also gives emphasis on the research programme in Chemistry. • The practical courses have been designed to prepare the students to have experience of the laboratory skills in Chemistry so that students can work in any scientific laboratories which are the need in the current scenario for becoming the ATMANIRBHAR. Students will be able to design and conduct experiments as well as to analyze and interpret scientific data in useful form. • Program will equip students to face the employment challenges and instill confidence to turn into entrepreneur and also step into research career. • The program will offer students with the knowledge and skill base that would enable them to undertake advanced studies in Chemistry and related areas or in multidisciplinary areas that involve Chemistry. • The students will gain domain knowledge and have right temperaments to know how to lead for successful career in academia, industry and research.

Program Specific Outcomes (PSOs):
• After successful completion of M.Sc. Chemistry program, the student will be able to create an awareness of the impact of chemistry on the society and development outside the scientific community. • Student will be able to analyze data obtained from various instruments viz. UV-Vis, Fluorescence, FTIR, NMR, TGA/DTA/DSC, GLC, GSC and HPLC for the structure determination and chemical analysis and student can apply different appropriate approach towards planning and execution of research in frontier areas of chemical sciences. • After successful completion of this program, student can apply different appropriate approach towards planning and execution of research in frontier areas of chemical sciences. • Student will become professionally trained and have caliber to do job in the various industries at all level of chemical, pharmaceutical, food products, life-oriented material industries.

Semester I Syllabus
Core Course
Course Code: B020701T
Course Title: Spectroscopy-I

Credit 05 **Hour 60**

Course Objectives: Students will be provided knowledge about fundamental concepts, tools and techniques used behind UV-visible, Infrared & Raman, microwave, mossbauer and diffraction techniques for structural determination of molecules.

Unit I

UV-Visible Spectroscopy: Different type of electronic transitions, Lambert-Beer's law, Chromophores, Auxochromes, Solvent effect, Red shift and blue shift, Woodward's rule for conjugated cyclic and acyclic dienes and α, β – unsaturated carbonyl compounds, Absorption in aromatic compounds (substituted benzene, naphthalene and anthracene), Problems related to UV-Visible Spectroscopy.

Unit II

Infrared Spectroscopy & Raman Spectroscopy: Linear harmonic oscillator, Vibrational energies of diatomic molecules, zero-point energy, force constant and bond strength, anharmonicity, Morse potential, Vibration– rotation spectroscopy, P, Q, R branches, Breakdown of Oppenheimer approximation, vibration of polyatomic molecules, selection rules, Group frequencies, Overtones, hot bands, factors affecting the bond positions and intensities for IR region, Problems related to Infrared Spectroscopy; *Raman Spectroscopy:* Classical and quantum theories of Raman effects, Pure rotational, Vibrational and Vibrational-rotational Raman spectra, Selection rule, Mutual exclusion principle, Resonance Raman spectroscopy, CARS.

Unit III

Microwave Spectroscopy: Classification of molecules, rigid rotor model, effect of isotopic substitution on the transition frequency, intensities, non-rigid rotor. Stark effect, nuclear and electron spin interaction and effect of external field applications.

Unit IV

Mossbauer Spectroscopy: Basic Principles, spectral parameters and spectrum display. Application of the technique to the studies of (a) bonding and structures of Fe^{+2} and Fe^{+3}compounds including those of intermediate spin, (b) Sn^{+2} and Sn^{+4} compounds – nature of M-L bond, coordination number, structure and (c) detection of oxidation state and in equivalent MB atoms.

Unit V

Diffraction Techniques: *X-ray Diffraction:* General Features of diffraction, Powder X-ray diffraction, Single crystal X–ray diffraction. The technique, structure factor, phase problem, brief description of time resolved X-rays diffraction techniques; *Electron Diffraction:* Scattering intensity vs scattering angle, Wierl equation, Measurement technique, Elucidation of structure of

simple gas phase molecules, Low energy electron diffraction structure of surfaces; *Neutron Diffraction:* Brief introduction, difference with X-rays diffractions.

Recommended Books:

1. Fundamentals of Molecular Spectroscopy, 4th Ed. Mc Graw-Hill, C.N. Banwell.
2. Basic Principles of spectroscopy, Mc Graw –Hill, R. Chang
3. Modern Spectroscopy, J. M. Hollas, John Wiley.
4. Inorganic Electronic Spectroscopy, A. B. P. Lever, Elsevier.
5. Magnetochemistry, R. L. Carlin, Springer Verlag
6. K. Veera Reddy, Symmetry and Spectroscopy of Molecules, New Age.
7. Inorganic Chemistry, D. E. Shriver, P. W. Atkins and C. H. L. Langford, Oxford
8. Advanced Inorganic Chemistry, F. A. Cotton and G. Wilkinson, John Wiley.
9. Inorganic Chemistry, J. E. Huheey, Ellen A. Keiter, Richard L. Keiter, Addison Wesley Longman (Singapore) Pvt. Ltd.

Course Outcomes:

CO-1. Students will acquire basic knowledge of UV-visible, infrared, Raman, microwave and diffraction techniques.

CO-2. Students will able to interpretate the spectra obtained from the various spectral (UV-visible, infrared & Raman, microwave, mossbauer and diffraction methods) techniques.

CO-3. Students will able to focus their aim for future prospects of research in the above these spectroscopic techniques.

Semester I Syllabus
Core Course
Course Code: B020702T
Course Title: Symmetry & Molecular Vibrations

Credit 05 **Hour 60**

Course Objective: Objective of this course is to provide deeper understanding about symmetry, points groups and relation between molecular vibrations and molecular structure which is strictly defined by the symmetry.

Unit I

Symmetry & Point Groups: Symmetry elements, Symmetry operations. Symmetry points group, Identification of molecular points group, Molecules of low symmetry, high symmetry and special symmetry (C_n, S_n, D_n, C_{nv}, D_{nh} only).

Unit II

Group, Subgroups, Classes and Matrices Representation: Definition, multiplication tables, group generating elements, subgroup, classes, derivation of matrices (C_n, σ, i, S_n), Direct product, Group multiplication basis, matrix representation, Character of an operation, orthogonality projection and shift operators, character table, orthogonality theorem, irreducible representation, Transformation matrices, structure of character table, determination of symmetry species for translations and rotations, construction of character table (C_{2v}, C_{3v}).

Unit III

Valence Bond Theory: Formation of hybrid orbitals of XY_3 (planar), XY_4. (tetrahedral & square planar), Symmetry of orbital, orbital symmetry properties, Projection to get symmetry orbital, projection operations, basis functions and hybrid orbitals with example.

Unit IV

Normal Coordinate Analysis: Cartesian coordinate and internal coordinate methods applied to C_{2v} (symmetric XY_2, ZXY_2), C_{3v} (XY_3), T_d. (XY_4) and O_h (XY_6) systems.

Unit V

Molecular Vibrations: Internal and symmetry coordinates, SALC's, Symmetric normal vibrations, mixing of linear coordinates in normal modes, determination of symmetry types of normal modes, analysis of vibration of 1, 2 dichloroethylene, IR and Raman activity of some typical molecules (C_{2v} C_{3v}, C_{4v}, D_{2h}, D_{3h}, D_{4h} point group).

Recommended Books:

1. McWeeny, "Symmetry - An Introduction to Group Theory", Pergamon Press.
2. Lowell H. Hall "Group Theory and Symmetry in Chemistry", McGraw Hill Book Company, New York, 1969
3. K. Veera Reddy, "Symmetry and Spectroscopy of Molecules", New Age International Limited Publisher, New Delhi.
4. D.M. Bishop, "Group theory and Chemistry" Dover Publications.
5. F.A. Cotton, "Chemical Applications of Group Theory", John Wiley, 1971.
6. M. Hamaresh, "Group theory and its Applications to Physical Problems" Addison- Wisley
7. Symmetry and Group theory: Some chemical applications, Ramashankar and Suresh Ameta, Himanshu Publications, Udaipur, Delhi.

Course Outcomes:

After successfully completion of this course:

CO-1: Students will have better understanding insight the symmetry elements and symmetry operations.

CO-2: Students will acquire the knowledge of symmetry of normal vibrations, determination of normal modes, mixing of internal coordinates and normal coordinate analysis of molecules.

CO-3: Students will focus their aim for future prospects of research in the field of symmetry & group theory.

Semester I Syllabus
Core Course
Course Code: B020703T
Course Title: Organic Reaction Mechanism

Credit 05 **Hour 60**

Course Objective: Objective of this course is to provide deeper understanding about the organic reaction mechanism related to the various types of organic reactions.

Unit I

Principle of Reaction Mechanism: Potential energy diagram, Transition states and intermediates, methods of determining reaction mechanism, Labelling and kinetics isotopic effect and its importance in the determination of reaction mechanism, Hammond's postulate, Curtin Hammett principle, structural effects on reactivity, Hammett equation and linear free energy relation (LFER), substituent and reaction constants, Taft equation.

Unit II

Substitution Reaction: *Aliphatic Nucleophilic Substitution*: SN^1, SN^2, SN^i, $SN^{1'}$, $SN^{2'}$,$SN^{i'}$, mixed SN^1 and SN^2, role of substrate's structure, nucleophile, leaving group and solvent on SN reaction, ambidentate nucleophile, Regioselectivity, competition between SN^1 and SN^2, Nucleophilic substitution in bridged system phenonium ion, norbornyl system, Neighbouring group participation (Ph, π, σ, N. S, negatively charged oxygen), anchimeric assistance; *Aliphatic Electrophilic Substitution*: SE^1 and SE^2, SE accompanied by double bond shifts, Effect of substrate, leaving group and solvent polarity on reactivity, *Aromatic Nucleophilic Substitution*: Aromatic SN^1 and SN^2, addition– elimination and elimination-addition (benzyne) mechanism, effect of substrate structure, nucleophile, leaving group on ArSN reaction, *Aromatic Electrophilic Substitution*: General view, energy profile, Arenium ion mechanism (ArSE), o/p ratio.

Unit III

Free Radical Reactions: Types, generation, structures, radical effect, substitution mechanism at an aromatic substrate at a bridgehead, reactivity in the attacking radicals, effect of solvent on reactivity, Allylic halogenation (NBS), oxidation of aldehydes, autooxidation, Alkynes coupling and arylation of aromatic compounds by diazonium salts, Sandmeyer reaction, Hunsdiecker reaction.

Unit IV

Elimination Reaction: E^1, E^2, E1cB, Factors (substrate structure, attacking base, leaving group, medium) affecting these reactions, stereochemistry, orientation of the double bond (Saytzeff vs Hofmann elimination), Mechanism and orientation of pyrolytic syn elimination, Competition between substitution and elimination reactions.

Unit V

Addition Reaction: *C=C bond addition*: Mechanism, stereochemistry, electrophilic, nucleophilic, free radical addition, addition of halogen acid, 1,2-dihydroxylation, epoxidation, hydroboration, oxymercuration-demercuration, hydrogenation of double bond, triple bond, aromatic ring. cyclopropanation, Simmon-Smith cyclopropanation, epoxidation, Sharpless asymmetric epoxidation, corey epoxidation, *Carbon hetero atom multiple bond addition*: C=O bonds, cram rule, condensation reactions involving Claisen, Benzoin, Perkin, Knoevenagel, Darzen, Reformatsky and Cannizzaro reaction, Mechanism of hydrolysis of ester and amide, Ammonolysis of ester.

Recommended Books:

1. Organic Chemistry, J. Clayden, N. Greeves, S. Warren and P. Wothers (Oxford Press.)
2. Advanced Organic Chemistry, A. F. A. Carey and R. J. Sundberg, 5th Ed. Springer (2007)
3. Advanced Organic Chemistry, J. March, 6th Ed.
4. Mechanism and structure in Organic Chemistry – E. S. Gould (Holt, Rinehart and Winston)
5. Guidebook to Mechanism in Organic Chemistry, Orient Longman, Sykes, P. A New Delhi.

Course Outcomes:

After successfully completion of this course:

CO-1: Students will train about the mechanistic approach about the mechanism of organic reactions.

CO-2: Students will acquire the knowledge general principle & outcome of various kinds of reactions viz. substitution reactions, addition reactions, elimination reactions and free radical reactions.

CO-3: Students will get idea to interpretate the mechanism of organic reactions and their products & prepare the students for further research in synthetic organic chemistry.

Semester I Syllabus
First Elective Course
Course Code: B020704T
Course Title: Chemical Kinetics & Thermodynamics

Credit 05 **Hour 60**

Course Objective: The objective of this course is to train the students about new and advance understanding of various concepts and applications of chemical kinetics and thermodynamics.

Unit I

Chemical Kinetics: Methods of determining rate laws, mechanism of photochemical ($H_2 + Br_2$, $H_2 + Cl_2$), chain stopped (H_2 equation + Br_2, decomposition of CH_3CHO, decomposition of C_6H_6), oscillatory reaction, collision theory, steric factor, Absolute reaction rate theory, comparison of result with Eyring and Arrhenius equation, steady state kinetics, kinetic and thermodynamic control of reactions, ionic reactions, kinetic salt effect, homogeneous catalysis, kinetics of enzyme reactions, heterogeneous catalysis, Fast reactions, luminescence and electron transfer process, flow technique, Relaxation method, flash photolysis, magnetic resonance method.

Unit II

Chemical Equilibrium: Free energy, entropy of mixing, partial molar quantities (free energy, volume, heat contents), Gibbs Duhem equation, Equilibrium constants, Van't Hoff equation, Fugacity and its determinations.

Unit III

Ideal and Non-ideal Solution: Excess function, activities, hydration number, activities in electrolytic solution, mean ionic activity coefficient, determination of activity, Phase rule, Phase diagram of one, two and three component systems.

Unit IV

Non-Equilibrium Thermodynamics: Postulates, methodologies, linear laws, Gibbs equation, Onsager reciprocal theory;

Unit V

Statistical Thermodynamics: Thermodynamic probability and entropy, concept of distribution, most probable distribution, Ensemble averaging, Maxwell-Boltzmann distribution, postulates of canonical, grand canonical, microcanonical ensemble, Bose-Einstein and Fermi-Dirac statics, partition function, translational, rotational, vibrational and electronic partition function for diatomic molecules, calculation of thermodynamic functions and equilibrium constant, theories of specific heat for solids, application of partition function.

Recommended Books:

1. P. W. Atkins, Physical Chemistry, Oxford University Press, New York.
2. S. Glasston, Physical Chemistry, Nostrand
3. K. L. Kapoor, Advance Physical Chemistry (Vol-1,2,3,4), Mac Millan, India
4. Puri, Sharma, Pathania, Advance Physical Chemistry.
5. M.C. Gupta, Statistical Thermodynamics, Second Edition, New Age International Limited Publisher.
6. Statistical Thermodynamics, Second Edition, New Age International Limited Publisher, India by M.C. Gupta.

Course Outcomes:

CO-1. Students will be able to understand the various terms and concepts behind the chemical kinetics and thermodynamics.

CO-2. Students will be able to explore the phase rules and phase diagram thermodynamically.

CO-3. Students will have knowledge of molecular orbital theory in term of charge calculation, equilibrium constants, excess function activities, Gibbs Helmholtz, Gibbs Duhem & Van't Hoff equation.

CO-4. Students will have knowledge of Non-Equilibrium Thermodynamics.

Semester I Syllabus
First Elective Course
Course Code: B020705T
Course Title: Surface & Solid-State Chemistry

Credit 05 **Hour 60**

Course Objectives: Objective of this course is to provide the detailed knowledge about new and advance understanding of various terms and concepts behind the surface chemistry, solid state reactions and electronic properties and band theory of the solids and their applications.

Unit I

Surface Chemistry: *Adsorption*: Surface Tension, capillary action, Laplace equation, Kelvinequation, Gibb's adsorption isotherm, BET equation, Electric kinetic phenomenon.

Unit II

Micelles: Surface active agents, classification, micellization, hydrophobic interaction, CMC, factors affecting CMC, Counter ion binding to micelles, solubilization, microemulsion, reverse micelles.

Unit III

Solid State Reactions: General principles, experimental procedures, co-precipitation as a precursor to solid state reactions, kinetics of solid-state reactions, Crystal Defects and Non-Stoichicmetry, Perfect and imperfect crystals, intrinsic and extrinsic defects – point defects, line and plane defects, vacancies, Schottky and Frenkel defects. Thermodynamics of Schottky and Frenkel defect formation, colour centers, non-stoichiometry and defects.

Unit IV

Electronic Properties and Band Theory: Metals, insulators and semiconductors, electronic structure of solids-band theory, band structure of metals, insulators and semiconductors. Intrinsic and extrinsic semiconductors, doping semiconductors, p-n junctions, super conductors. Optical Properties – Optical reflectance, photoconduction.

Unit V

Magnetic Properties: Classification of materials: Quantum theory of paramagnetic- cooperative phenomenal magnetic domains, hysteresis; *Organic Solids*: Electrically conducting solids, organic charge transfer complex, organic metals, new superconductors.

Recommended Books:

1. Solid Sate Chemistry and its Application, A. R. West, Plenum.
2. Principal of The Solid state, H. V. Keer, Wiley Eastern.
3. Solid Sate Chemistry, N. B. Hannay.
4. Solid Sate Chemistry, D.K. Chakrabarty, New age International.

Course Outcomes:

CO-1. Students will be able to understand the basics of surface chemistry, solid state reactions and electronic properties and band theory of the solids.
CO-2. Students will be able to apply the theoretical concepts of this course in proper applications.
CO-3. This course will motivate students to enhance their knowledge in this field via research.

Semester I Syllabus
Second Elective Course
Course Code: B020706P
Course Title: Chemistry Laboratory Course-I A

Credit 05 **10h/Week**

Course Objectives: The objective of this course to provide advanced insight about the qualitative analysis of inorganic mixture containing one rare element of first & second group and organic mixtures, chromatographic separations, chemical kinetics, thermochemistry and phase equilibria.

INORGANIC CHEMISTRY

Qualitative analysis: Qualitative analysis of an inorganic mixture of seven radicals including one rare element of first & second group (TI, W, Se, Mo, and Te). Semi- micro analysis is to be done. Mixture can have insoluble substances, interfering anions and combination of anions.

Chromatography: Chromatographic separation of first and second group metal ion of the following combinations:
i. Pb^{2+}, Ag^{+}, Hg_2^{2+}
ii. Pb^{2+}, Cd^{2+}, Cu^{2+}
iii. Bi^{3+}, Cd^{2+}, Hg^{2+}

ORGANIC CHEMISTRY

Qualitative Analysis: Separation, purification and identification of components of binary organic mixture (both solids, one solid & one liquid). Systematic analysis of each component laying emphasis on solubility, element detection, melting point, boiling point determination, ignition test, unsaturation test, functional group test, specific test and preparation of suitable derivative.

Chromatography: Separation, identification and determination of Rf value of the components present in the binary mixture of amino acids by paper chromatographic methods.

PHYSICAL CHEMISTRY

Chemical Kinetics:

i. Kinetics of ester (methyl acetate) hydrolysis in presence of acid.
ii. Determine the velocity constant and order of reaction for hydrolysis of ethyl acetate by sodium hydroxide at given temperature (saponification of an ester).
iii. Kinetics of acetone and I_2
iv. Kinetics of $KBrO_3$/ KI
v. Kinetics of $Na_2S_2O_8$ / HCl

Thermochemistry/Phase Equilibria:

i. Determination of the solubility of benzoic acid in water at different temperatures and calculate the heat of solution.
ii. Determination of the solubility of a salt (KCl, KNO_3) at different temperatures and calculate the heat of solution.

iii. Construct the phase diagram for three component Ethanol, benzene and water system.

iv. Construct the phase diagram for three component chloroform, acetic acid and water system.
v. Construct the phase diagram for two component system.

System of Marking:-
Time: 12h
Inorganic: 33
Organic: 34
Physical: 33

Recommended Books:

1. Experimental Inorganic Chemistry by Mounir A, Malati, Horwood series in Chemical Science (Horwood publishing Chichester) 1999.
2. Practical Inorganic Chemistry, G. Marrand, B.W. Rockett, Van Nostrand.
3. Practical Physical Chemistry, A.M. James and F.E. Prichand, Longman.
4. Findley's Practical Physical Chemistry revised, B.P. Levitt, Longman.
5. Experimental Physical Chemistry, R.C. Das and Bebera, Tata Mc Grawhill.
6. Senior Practical Physical Chemistry, B.D. Khosla and V.S. Barg (R. Chand and Co., Delhi)
7. Experimental Physical Chemistry by D.P. Shoemaker Mc Grawhill, 7th Edition 2003.
8. Experiments in Chemistry, D.V. Jahagirdar, Himalaya Publishing House.
9. Practical Physical Chemistry, B. Vishwanathan and P.S. Raghwan, Viva Books.
10. General Chemistry Experiments, Anil J Elias, University Press (2002)
11. Experimental Physical Chemistry, V.D. Athawale, ParulMathur, New Age International (P) Limited.
12. Systematic Experiment in chemistry, ArunSethi, New Age International (P) Limited.
13. The systematic Indentification of Organic Compounds, R.L. Shringer and D.Y. Curlin.

Course outcomes:

After completion of this course students will acquire the knowledge of:
CO-1. Qualitative analysis of inorganic mixtures containing one rare element of first & second group & paper chromatographic separation techniques of cations.
CO-2. Qualitative analysis of binary organic mixture and paper chromatographic technique for the separation of mixture of amino acids.
CO-3. Chemical kinetics, thermochemistry and phase equilibria.

Semester I Syllabus
Second Elective Course
Course Code: B020707P
Course Title: Chemistry Laboratory Course-I B

Credit 05 **10h/Week**

Course Objectives: The objective of this course to provide advanced insight about the qualitative analysis of inorganic mixture containing one rare element of third group and organic mixtures, paper/thin layer chromatographic technique, colloidal & surface chemistry and molecular weight measurements.

INORGANIC CHEMISTRY

Qualitative analysis: Qualitative analysis of an inorganic mixture of seven radicals including one rare element of third group (Be, U, Ti, Zr, Th, Ce, V and Li). Semi- micro analysis is to be done. Mixture can have insoluble substances, interfering anions and combination of anions.

Chromatography: Chromatographic separation of third, fourth & fifth group metal ion of the following combinations:
i. Fe^{3+}, Cr^{3+}, Al^{3+}
ii. Co^{2+}, Ni^{2+}, Zn^{2+}
iii. Ba^{2+}, Sr^{2+}, Ca^{2+}

ORGANIC CHEMISTRY

Qualitative Analysis: Separation, purification and identification of components of binary organic mixture (both liquids), Systematic analysis of each component leading to their final identification laying emphasis on solubility, element detection, melting point, boiling point determination, ignition test, unsaturation test, functional group test, specific test and preparation of suitable derivative.

Chromatography: Separation, identification and determination of Rf value of the components present in the binary mixture of sugars by thin layer chromatographic methods.

PHYSICAL CHEMISTRY

Colloidal & Surface Chemistry:

i. Critical Micelle Concentration (CMC) of Surfactants by surface tension measurement,
ii. To determine the adsorption of aq. acetic acid by activated charcoal and study the adsorption (Freundlich isotherm).
iii. To study the adsorption of oxalic acid on charcoal and draw the Freundlich isotherm.
iv. Find out surface tension, molecular energy and Parachor of given liquid at room temperature.
v. To prepare As_2O_3 and $Fe(OH)_3$ sol.
vi. To prepare $Al(OH)_3$ sol and sulphur sol.

Molecular Weight Measurements:

i. To determine molecular weight of a given electrolyte by elevation in boiling point method (Landsberger method) and also find out its Van't Hoff factor.
ii. To determine molecular weight of a given polymer by viscosity method.
iii. Determine molecular weight of a given electrolyte by depression in freezing point method.

System of Marking:-
Time: 12h
Inorganic: 33
Organic: 34
Physical: 33

Recommended Books:

1. Experimental Inorganic Chemistry by Mounir A, Malati, Horwood series in Chemical Science (Horwood publishing Chichester) 1999.
2. Practical Inorganic Chemistry, G. Marrand, B.W. Rockett, Van Nostrand.
3. The systematic Indentification of Organic Compounds, R.L. Shringer and D.Y. Curlin.
4. Practical Physical Chemistry, A.M. James and F.E. Prichand, Longman.
5. Findley's Practical Physical Chemistry revised, B.P. Levitt, Longman.
6. Experimental Physical Chemistry, R.C. Das and Bebera, Tata Mc Grawhill.
7. Senior Practical Physical Chemistry, B.D. Khosla and V.S. Barg (R. Chand and Co., Delhi)
8. Experimental Physical Chemistry by D.P. Shoemaker Mc Grawhill, 7th Edition 2003.
9. Experiments in Chemistry, D.V. Jahagirdar, Himalaya Publishing House.
10. Practical Physical Chemistry, B. Vishwanathan and P.S. Raghwan, Viva Books.
11. General Chemistry Experiments, Anil J Elias, University Press (2002)
12. Experimental Physical Chemistry, V.D. Athawale, ParulMathur, New Age International (P) Limited.
13. The systematic Indentification of Organic Compounds, R.L. Shringer and D.Y. Curlin.
14. Systematic Experiment in chemistry, ArunSethi, New Age International (P) Limited.

Course outcomes:

After completion of this course students will acquire the knowledge of:
CO-1. Qualitative analysis of inorganic mixtures having one rare element of third group & paper chromatographic separation techniques of metal ions of third group.
CO-2. Qualitative analysis of binary organic mixture having both liquid components and TLC separation of mixture of sugars.
CO-3. Colloidal & surface chemistry and molecular weight measurements.

Semester II Syllabus
Core Course
Course Code: B020801T
Course Title: Chemistry of Main Group Elements

Credit 05 **Hour 60**

Course Objectives: Objective of this course is to prepare the students to understand and correlate preparation, structure, bonding and properties of main block elements.

Unit I

Stereochemistry of Bonding Among Main Group Elements: VSEPR theory stereochemical rules and explanation of the shapes of molecules and ions of nontransition element with 2-7 valence shell electron pairs. Walsh diagram (Tri and penta atomic molecules) $d\pi$-$p\pi$ bonds, Bent rule, Energetics of hybridization.

Unit-II

Compounds of Main Group Elements: Preparation, Structure, Bonding and Technical Applications of Polyether complexes of alkali and alkaline earth metals; Polyphasphazenes and Thiazyl & its polymers, tetrasulfur dinitride.

Unit-III

Structure and Bonding in Ions of Some Main Group Elements: Structure and bonding of borane anions, higher boranes, carboranes, classification and structures of silicates.

Unit-IV

Carbides & Polyanions: Synthesis and structure of Carbides & polyions of Ge, Sn, Pb, Sb, Bi and Hg.

Unit-V

Hapticity & Organometallics: Definition and classification of organometallic compounds on the basis of hapticity and polarity of metal-carbon bond; Preparation, Properties, Structure and Applications of alkyl and aryls of Lithium, Beryllium, Aluminum, Mercury and Tin.

Recommended Books:

1. Advance Inorganic Chemistry, 6th Edition, Cotton and Wilkinson.
2. Inorganic Chemistry, 4th Edition, Principles of Structure and Reactivity by J.F. Huheey, E.A. Keiter and R.L. Keiter, 1993
3. Chemistry of Elements by N.N. Greenwood and A. Ernshaw, Butterworths 1997.

4. Organometallic Chemistry: A Unified Approach by R.C. Mehrotra and A.K. Singh
5. Comprehensive Coordination Chemistry Vol.3 by G. Wilkinson, R.D. Gillard, And J.A. McCleverty, Pergamon Press 1987.

Course Outcomes:

After completion of this course students will be able to:

CO 1. Understand of correlation between electronic configuration and bonding properties of main group elements.

CO 2. Focus their aim for future prospects of research in the field of chemistry of main group elements.

Semester II Syllabus
Core Course
Course Code: B020802T
Course Title: Stereochemistry & Spectroscopy-II

Credit 05 **Hour 60**

Course Objectives: Objectives of this course is to provide detailed knowledge about the relative special arrangement of atoms that form the structure of molecules and their manipulation. Knowledge of spectroscopy is also needed as it helps to elucidate the structure of the compound.

Unit I

Stereochemistry: Stereochemistry with chiral centre: chirality, Polychiral centre molecules, Threo-erythro isomers, stereoisomerism with axial/ planar chirality and Helicity, Principle of axial and planar chirality, optical isomerism in the absence of chiral carbon (biphenyl, allenes, spiranes), optical activity due to intermolecular overcrowding, chirality due to helical shape, Absolute configuration (R/S, E/Z), Stereochemistry of compounds having S, N, P atoms, Geometrical isomerism of compounds having C=N, N=N bonds; *Topocity and Prostereoisomeriam*: Homotopic, enantiotopic and diastereotopic atoms, groups and faces, nomenclature and symbols; *Atropisomerism*: Conformational analysis of acyclic system, Interconversion of Fischer, Newmann and Sawhorse projection, its effect on activity (SN^1, SN^2, E^1, E^2) configuration, conformation and stability of cycloalkanes, mono and disubstituted cyclohexane, cyclohexenones, decalin, decalol.

Unit II

Proton Magnetic Resonance Spectroscopy: Spinning nuclei, nuclear spin, nuclear resonance, saturation, chemical shift, chemical shift measurement, factors affecting the chemical shift, anisotropic effect, shielding mechanism, spin - spin coupling, coupling constant, chemical exchange, effect of deuteration, factor influencing coupling constant 'J', Karplus curve-variation of coupling constant with dihedral angle, Spin decoupling, simple, virtual and complex coupling, chemical and magnetic equivalence, first and non -first order spectra, analysis of AB, AMX and ABX system, simplification of complex spectra, contact shift reagents, solvent effects. NOE, hindered rotation and rate process, NMR studies of ^{19}F, ^{31}P, instrumentation, FT NMR & its advantages, DEPT, 2DNMR: COSY, NOESY, HETCOR, application of ^1H NMR spectra in structural determination of simple organic molecules, use of NMR in medical.

Unit III

Carbon Magnetic Resonance Spectroscopy: Introduction, peak assignment, chemical shift, ^{13}C-^1H coupling, off resonance, decoupling, deuterium, fluorine and phosphorous coupling, DEPT, 2DNMR: COSY, NOESY, HETCOR, application to simple organic molecules (aliphatic, olefinic, alkyne, aromatic, heteroaromatic and carbonyl carbon).

Unit IV

ESR Spectroscopy: Basic principles, zero field splitting and Kramer's degeneracy, factors affecting 'g' value, Isotropic and anisotropic hyperfine coupling constants, Application to organic free radicals - methyl free radical, naphthalene and benzene free radicals, CIDNP.

Unit V

Mass Spectrometry: Measurement technique (EI, CI, FD and FAB), Ion production, factors affecting fragmentation, group metastable peak, Ion analysis, Molecular base and molecular ion, Ion abundance, factors affecting Ion abundance, Mc Lafferty rearrangements, retro Diels Alder fragmentation, Nitrogen rule, determination of molecular composition, fragmentation patterns of organic compounds, common functional group with reference to their structure determination, Interpretation of mass spectra, High resolution mass spectrometry.

Recommended Books:

1. Stereochemistry of Organic Compounds, Nasipuri, New Age International (P) Limited.
2. Stereochemistry of Carbon Compounds, E. L. Eliel and S. H. Wilen
3. Spectrometric Identification of Organic Compounds, Silverstein and Webster, John Wiley, New York.
4. Organic Spectroscopy, P. S. Kalsi, New Age International (P) Limited.
5. Introduction to Spectroscopy, Pavia, Lampman, Kriz, Vyvyan, Cengage Learning.
6. Organic Spectroscopy, I Fleming, McGraw-Hill Inc., US.
7. Organic Spectroscopy, W. Kemp, Macmillan, London.

Course Outcomes:

CO-1. Students will be able to demonstrate an intuitive understanding of concepts of stereochemistry and spectroscopy I.

CO-2. After studying this course, the students will be able to have understanding of various classes of stereoisomers and also have knowledge to establish the structure of the molecules by analyzing their various spectral data.

CO-3. Stereochemistry and spectroscopy I cover a wide area of research in organic chemistry and hence this course will motivate students to enhance their knowledge in this field via research.

Semester II Syllabus
Core Course
Course Code: B020803T
Course Title: Advanced Quantum Mechanics

Credit 05 **Hour 60**

Course Objectives: Objective of this course to provide detailed knowledge to the students about molecular structures and properties, quantum states live in a vector space, to relate abstract formulation to wave and matrix mechanics, perturbation theory, level splitting and radiative transitions, role of angular momentum in atomic and nuclear physics, relation between conservation laws and symmetries.

Unit I

Symmetry Properties and Quantum Mechanics: Invariability of Schrodinger Equation for a molecule with respect to symmetry operations and its consequences. Construction of molecular orbitals of ammonia and pie molecular orbitals of naphthalene, the direct product representation and its application in the derivation of selection rules for electronic, vibrational and Raman spectra.

Unit II

Huckel MOT of Conjugated Systems and its Applications: Calculation of energy levels and delocalization energy of butadiene, cyclic conjugated systems: cyclopropenyl, cyclobutadiene, cyclopentadienyl, benzene, brief idea about delocalization energies of tropylium radical and cyclooctatetraene, concept of aromaticity and antiaromaticity, Huckel treatment of linear polyenes.

Unit-III

Semi-Empirical and Ab-Intio SCF Theories: Hartee-Fock Self consistent (SCF) method, Semi empirical SCF theory (CDNO, INDO & MNDO), Slater and Gaussian type orbitals, Configurational interaction and electron correlation, Moellar-Plasset perturbation methods.

Unit-IV

Introduction to Density Functional Theory: Concept of basis sets, exchange- correlation energy, The Hohenberg variational theorem and Kohn- Sham orbitals, The Local Density Approximation (LDA) and Generalized Gradient Approximation (GGA). Density Functional theory and its significance.

Unit-V

Molecular Mechanics: A brief introduction to molecular mechanics.

Recommended Books:

1. Quantum Chemistry by Donald A. Macquarrie
2. Physical Chemistry by T. Engel and P. Reid
3. Introductory Quantum Chemistry by A. K. Chandra
4. Quantum Chemistry by R. K. Prasad
5. Molecular Quantum Mechanics by Atkins and Friedman
6. Quantum Chemistry by Ira N. Levine, Prentice Hall of India, New Delhi 1995
7. Chemical Application of Group Theory by F. A. Cotton

Course Outcomes:

After completion of this course students will able to:

CO 1. Understand about molecular structures and properties, quantum states live in a vector space, to relate abstract formulation to wave and matrix mechanics.

CO-2. Understand about perturbation theory, level splitting and radiative transitions, role of angular momentum in atomic and nuclear physics relation between conservation laws and symmetries.

CO-3. Focus their aim for future prospects of research in the field of quantum chemistry.

Semester II Syllabus
Third Elective Course
Course Code: B020804T
Course Title: Research Aptitude

Credit 05 **Hour 60**

Course Objectives: The objective of this course is to create the interest of the students by providing the knowledge of research methodology in order to carry out the good quality of research work and writing the scientific articles systematically.

Unit I

Elementary of Basic Research: Definition & objectives of research, motivation in research, research, methodology, literature survey, primary source, secondary source, reprint, importance of literature survey in defining problem and selection of topic of research, distinction between theoretical and applied research, various stages of scientific research, criteria of good research, research area in chemical sciences, significance of research in chemical sciences.

Unit II

Chemical Abstract and Review Writing: Chemical abstract types, survey of abstract indexes (substance index, author index, general technique index, collective and comprehensive indices); monograph and treaties on specific areas, understanding of terminology text, reference, comprehensive survey, vocabulary regarding book, citation index, impact factor, H-index, e-consortium, preparation of a review article related to the research problem, writing scientific report, planning of writing, preparation, draft, revision, refining, research report format, writing literature surveys and reviews.

Unit III

Digital Resources for Research and Error Analysis: Web resources, e-Journal access, UGC infonet, e-books, internet discussion, groups and communities blog, preprint server, search engines-scirus, google scholar, chemindustry, wiki-database, chemspider, science direct, sciFinder, scopus, elementary of error's analysis.

Unit IV

Types of Scientific Papers and its Communications: Original research papers, patents, review papers, symposium papers, invited papers, conference proceeding as full papers, format of research papers, components of thesis; *Special Elements*: Footnote, number, quantities, SI units, functions, mathematical expression and equations, tables, figure, captions, link between figure and text, line drawing, diagrams and graphs, punctuation, common proof marks that may be used to correct a manuscript; *Research Work Writing & Presentations:* poster display, oral presentation, knowledge about seminar, symposium, conferences, convention, congress, workshop.

Unit V

Financial Assistance & Research Ethics: Elementary of writing of research proposal, role of funding agencies like UGC, CSIR, ICAR, ICMR, ISRO, DRDO, DST, DBT, CST in R&D of chemical sciences; *Plagiarism*: Definition, consequences, avoidance.

Recommended Books:

1. Research Methodology, C.R Kothari, New Age International Publication, 2004.
2. J Writing and Presentation Scientific Papers 2 /e; Malmfors, Grossman, Viva Book Pvt. Ltd.
3. How to write a successful science Thesis; Russey, Ebel, Bliefest, Wiley-VCH.

Course Outcomes:

After completion of this course students will be able to:
CO 1. Understand the basic of this course and think & develop new ideas in this course.
CO 2. Understand how-to do-good quality of research work systematically.
CO 3. Understand how to do write the research projects and scientific papers.

Semester II Syllabus
Third Elective Course
Course Code: B020805T
Course Title: Environmental Science

Credit 05 **Hour 60**

Course Objectives: The objective of this course to provide knowledge about the basics of environment, atmosphere, hydrosphere, industrial pollution and chemical toxicology and method of analysis of pollutants.

Unit I

Basics of Environment: Introduction, Composition of atmosphere, Vertical temperature, heat budget of the earth atmospheric system, vertical stability atmosphere. Biogeochemical cycles of C, N, P, S and O, bio distribution of elements.

Unit II

Atmosphere: Regions, chemical composition of atmosphere, , Particles, ions & radicals and their formation, Chemical & photochemical reaction in atmosphere, smog formation, oxides of N,C,S,O and their effect, pollution by chemicals, petroleum, minerals, ozone layer, ozone layer depletion, chlorofluoro hydrocarbons, Green House effect, acid rain, Global warming, Analytical methods for measuring air pollutants, continuous monitoring instruments.

Unit III

Hydrosphere: Chemical composition of water bodies *viz.* lakes, streams, rivers, and wet lands, Hydrological cycle, Recycle of waste water, Sewage treatment; *Water pollution*: Inorganic, Organic, Pesticide, Agricultural, Industrial and Sewage, detergents, oil spills and oil pollutants, Water quality parameters-dissolved oxygen, biochemical oxygen demands, solids metals, content of chloride, sulphate, phosphate, nitrate and micro-organisms water quality standards, Analytical methods for measuring DO, BOD and COD, residual chloride and chlorine demand, Purification and treatment of water.

Unit IV

Soil: Composition, micro and macro nutrients, Pollution- fertilizers, pesticides, plastics and metals, waste treatment.

Unit V

Industrial Pollution: Cement, sugar, distillery, drug, paper and pulp, thermal power plants, nuclear power plants, metallurgy, polymers, drugs etc. Radionuclide analysis, Disposal of Wastes and their

management; *Environmental Toxicology:* Toxic chemicals in the environment, Chemical sclution to environment, Principles of decomposition, biodegradability, better industrial process.

Recommended Books:

1. Manahan, Stanley E. Fundamentals of Environmental Chemistry Boca Raton: CRC Press, LLC, 2001.
2. Sonja Krause, Herbert M. Clark, James P. Ferris, Robert L. Strong Chemistry of the Environment, Elsevier Science & Technology Books, 2002.
3. Eugene R. Weiner Applications of Environmental Chemistry 2000 CRC Press, LLC.
4. By Clair, N. Sawyer, Perry L. Mc Carty, Gene F. Parking Chemistry for environmental engineering and Science (5^{th} edition) McGraw Hill Professional.

Course Outcomes:

After completion of this course:

CO-1. Students will be able to understand various terms and concepts behind the environmental chemistry.

CO-2. Students will have deeper understanding on various environmental issues viz. fate of chemical species in soil, water and air.

CO-3. Students will get deeper insight about industrial pollution and chemical toxicology.

Semester II Syllabus
Fourth Elective Course
Course Code: B020806T
Course Title: Chemistry Laboratory Course-II A

Credits-5 **10h/Week**

Course Objectives: The objective of this course to provide advanced insight about the quantitative estimations (one volumetrically and other gravimetrically) of the metals present in inorganic mixtures, inorganic and organic preparations, chromatographic separations of sugars, and experiments related to physical chemistry viz. Solubility & distribution coefficient, optical methods (colorimetry, refractometry and polarimetry).

INORGANIC CHEMISTRY

Quantitative Analysis: Estimation of two metal ions (one gravimetric and other volumetric) from the following mixtures:

i. Ni^{2+} and Cu^{2+}
ii. Pb^{2+} and Cu^{2+}
iii. Ba^{2+} and Cu^{2+}
iv. Ag^{+} and Cu^{2+}

Inorganic Preparation:
Check the purity of the synthesized compounds by TLC and report the percentage of yield.

i. $K_3[Cr(C_2O_4)_3].\ 3H_2O$
ii. $[Co\ (NH_3)_4.CO_3]NO_3$
iii. $NH_4\ [Cr\ (SCN)_4\ (NH_3)_2]H_2O$
iv. $(NH_4)_2\ [PbCl_6]$
v. $Hg\ [Co\ (SCN)_4]$
vi. $Ni(DMG)_2$

ORGANIC CHEMISTRY

Organic Synthesis:
Two steps synthesis involving following reactions and check the purity of the synthesized compounds by TLC and report the percentage of yield:-
i. Acetylation

ii. Hydrolysis
iii. Oxidation
iv. Aromatic electrophilic substitution
v. Condensation
vi. Sandmeyer reaction

Chromatography: Separation and identification of the sugars present in the organic mixture by paper chromatographic methods and determination of Rf value.

PHYSICAL CHEMISTRY

Solubility & Distribution Coefficient:

i. To draw the solubility curve for water-acetic acid-chloroform system.
ii. Determination of the distribution coefficient of acetic acid between benzene and water.
iii. Determination of the distribution coefficient of iodine between carbon tetrachloride and water.
iv. Determination of the dimerization constant of benzoic acid in benzene medium by partition method.
v. Solubility of an organic acid in water at room temperature.

Optical Methods (Colorimetry, Refractometry and Polarimetry):

i. To verify Lambert's Beer Law colorimetrically.
ii. Determine the rate constant for inversion of cane sugar using a polarimeter.
iii. Determination of the molar refractivity of methyl alcohol, acetic acid, ethyl acetate and carbon tetrachloride and calculate the refraction equivalents of carbon, hydrogen and chloride atoms.

System of Marking:-
Time: 12h
Inorganic: 33
Organic: 34
Physical: 33
Recommended Books:

1. Vogels Text book of Quantitative Analysis revised, J. Bessett, R.C. Denney, G.H. Jellery and J. Mendhan ELBS.
2. Microscale Inorganic Chemistry, Z. Scafran, R.M. Pike and M.M. Singh Wiley.
3. Practical Inorganic Chemistry, G. Marrand, B.W. Rockett, Van Nostrand.
4. Vogel's Textbook of Practical Organic Chemistry, A.R. Tatchell, John Wiley.
5. Systematic Experiment in chemistry, Arun Sethi, New Age International (P) Limited.
6. Practical Physical Chemistry, A.M. James and F.E. Prichand, Longman.
7. Findley's Practical Physical Chemistry revised, B.P. Levitt, Longman.
8. Experimental Physical Chemistry, R.C. Das and Bebera, Tata Mc Grawhill.

9. Senior Practical Physical Chemistry, B.D. Khosla and V.S. Barg (R. Chand and Co., Delhi)
10. Experimental Physical Chemistry by D.P. Shoemaker Mc Grawhill, 7th Edition 2003.
11. Experiments in Chemistry, D.V. Jahagirdar, Himalaya Publishing House.
12. Practical Physical Chemistry, B. Vishwanathan and P.S. Raghwan, Viva Books.
13. General Chemistry Experiments, Anil J Elias, University Press (2002)
14. Experimental Physical ChemistryPractical Physical Chemistry revised, B.P. Levitt, Longman.
15. Experiments in Physical chemistry, J.C. Ghosh, BharatiBhavan.
16. Advanced Practical Physical Chemistry, JB Yadav.

Course outcomes:

After completion of this course students will acquire the knowledge of:

CO-1. Quantitative estimation (one volumetrically and other gravimetrically) of the metals present in the inorganic mixtures.

CO-2. Paper chromatographic (paper) separations of sugars.

CO-3. Preparation of the inorganic and organic compounds

CO-4. Various theories and principles behind the experiment of . Solubility & distribution coefficient, optical methods (colorimetry, refractometry and polarimetry).

Semester II Syllabus
Fourth Elective Course
Course Code: B020807T
Course Title: Chemistry Laboratory Course-II B

Credits-5 **10h/Week**

Course Objectives: The objective of this course to provide advanced insight about the quantitative estimations (both gravimetrically) of the metals present in inorganic mixtures, chromatographic separations, inorganic and organic preparations as well as identification of synthesized compounds by using various spectral techniques and experiments related to physical chemistry viz. conductance measurement and flame photometry.

INORGANIC CHEMISTRY

Quantitative Analysis: Gravimetric estimation of two metal ions from the following mixtures:

i. Cu^{2+} and Ni^{2+}
ii. Cu^{2+} and Zn^{2+}
iii. Zn^{2+} and Ni^{2+}
iv. Cu^{2+} and Fe^{2+}
v. Ni^{2+} and Fe^{2+}
cc

Inorganic Preparation:

Identification of the synthesized inorganic compounds by using their spectral data (UV, IR and Mass spectroscopy)

i. Trans $K[Cr(C_2O_4)_2(H_2O)_2].2H_2O$
ii. Ferric alum (ferric ammonium sulphate)
iii. Prussian blue
iv. $[Cu(NH_3)_4]SO_4.H_2O$

ORGANIC CHEMISTRY

Organic Synthesis: Two steps synthesis
Identification of the synthesized organic compounds by using their spectral data (UV, IR, ^{1}HNMR, ^{13} CNMR and Mass spectroscopy).

Chromatography: Separation and identification of the components present in the given organic mixture by thin layer chromatographic methods and determination of Rf value.

PHYSICAL CHEMISTRY

Conductance Measurement:

i. Study the hydrolysis of aniline hydrochloride by conductance method.
ii. Determination of basicity of a given salt by conductance method.
iii. Conductometric titration of strong acid with strong base.
iv. Conductometric titration of weak acid with strong base.
v. Verification of Ostwald's dilution law.
vi. Verification of Kohlrausch's law

Flame Photometry:

i. Estimation of sodium and potassium in admixture.
ii. Estimation of magnesium and calcium in tap water.

System of Marking:-
Time: 12h
Inorganic: 33
Organic: 34
Physical: 33

Recommended Books:

1. Vogels Text book of Quantitative Analysis revised, J. Bessett, R.C. Denney, G.H. Jellery and J. Mendhan ELBS.
2. Microscale Inorganic Chemistry, Z. Scafran, R.M. Pike and M.M. Singh Wiley.
3. Practical Inorganic Chemistry, G. Marrand, B.W. Rockett, Van Nostrand.
4. Vogel's Textbook of Practical Organic Chemistry, A.R. Tatchell, John Wiley.
5. Systematic Experiment in chemistry, Arun Sethi, New Age International (P) Limited.
6. Practical Physical Chemistry, A.M. James and F.E. Prichand, Longman.
7. Findley's Practical Physical Chemistry revised, B.P. Levitt, Longman.
8. Experimental Physical Chemistry, R.C. Das and Bebera, Tata Mc Grawhill.
9. Senior Practical Physical Chemistry, B.D. Khosla and V.S. Barg (R. Chand and Co., Delhi)
10. Experimental Physical Chemistry by D.P. Shoemaker Mc Grawhill, 7th Edition 2003.
11. Experiments in Chemistry, D.V. Jahagirdar, Himalaya Publishing House.
12. Practical Physical Chemistry, B. Vishwanathan and P.S. Raghwan, Viva Books.
13. General Chemistry Experiments, Anil J Elias, University Press (2002)
14. Experimental Physical ChemistryPractical Physical Chemistry revised, B.P. Levitt, Longman.
15. Experiments in Physical chemistry, J.C. Ghosh, BharatiBhavan.

16. Advanced Practical Physical Chemistry, JB Yadav.

Course outcomes:

After completion of this course students will acquire the knowledge of:

CO-1. Quantitate estimation of the metals present in the inorganic mixtures.

CO-2. Thin layer chromatographic separations of component present in the organic mixture.

CO-3. Preparation and identifications of the inorganic and organic compounds by using various spectral techniques

CO-4. Various theories and principles behind the experiment of conductance measurement and flame photometry.

Semester III Syllabus
Core Course
Course Code: B020901T
Course Title: Coordination & Bioinorganic Chemistry

Credit 05 **Hour 60**

Course Objectives: After successful completion of the first year of Masters, students coming in III[rd] semester are to provide knowledge about energy level in an atom, free ions in crystal field, electronic spectra, metal ligand bonding & magnetic properties in transition metal complexes and also insight about bioinorganic chemistry.

Unit I

Basics of Energy Level in Atoms: Electronic configuration, coupling of orbital angular momenta, coupling of spin angular momenta, spin orbital coupling, energy terms, Determining the ground state terms -Hund's rules. Determination of term symbol for a closed subshell, Hole formulation, Derivation of terms for a d^2 configuration. Inter electron repulsion parameter. Variation of Racah B and C parameters in different transition series. Spin orbit coupling parameters.

Unit-II

Free Ions in Crystal Fields: Effect of weak crystal field on free ion terms in octahedral, square planar and tetrahedral symmetries. Orgel diagrams, mixing of terms, Medium and strong field approximation in Oh point group, transition from weak to strong field and correlation diagram for only d2 case, non-crossing rule, Tanabe Sugano diagrams.

Unit-III

Electronic Spectra of Complexes: Laporte orbital, selection rules, spin selection rules, Splitting of electronic energy level and spectroscopic states, Interpretation of the spectra of $[M(H_2O)_6]^{n+}$ in aqueous medium, calculation of Dq, B and β parameters, Jahn Teller distortion and its effect on electronic spectra.

Unit-IV

Metal-Ligand Bonding & Magnetic Properties of Complexes: Effects of crystal field splitting, Limitations of CFT, Nephelauxetic series, molecular orbital energy level diagram of octahedral, tetrahedral and square planer complexes; *Magnetic Properties of Complexes:* Dia, para, ferro and anti-ferromagnetism, Quenching of orbital angular momentum by ligand and magnetic properties of A, E and T terms.

Unit-V

Metalloenzymes: Function, structure, bonding and stereochemistry of the active site of *Natural dioxygen carriers*: Haemoglobin, myoglobin, hemerythrin, hemocyanin; *Electron Transport:* Iron sulphur protein- Rubredoxin, Ferredoxin, Cytochromes (types a, band c); *Redox enzymes*: Mo

containing: Nitrogenase, Xanthine oxidase, Sulphite oxidase, Nitrate reductase, Fe containing: Cyt c oxidase, Catalases Peroxidases; Cu containing: Superoxide dismutase (SOD), Bovine superoxide dismutase (BOD), Ascorbic acid oxidase, Zn containing: Carboxypeptidase A & B, Carbonic anhydrase, Urease, Co containing: Vitamin B_{12}, Vitamin B_1, Methyl cobalamime, Biomethylation.

Recommended Books:

1. Advanced Inorganic Chemistry, F. A. Cotton and G. Wilkinson, John Wiley
2. Inorganic Chemistry, J. E. Huheey, Ellen A. Keiter, Richard L. Keiter, Addison Wesley Longman (Singapore) Pvt. Ltd.
3. Chemistry of the Elements, N. N. Greenwood and A. Earnshow, Pergamon.
4. Organometallic Chemistry: A Unified Approach, R. C. Mehrotra and A. K. Singh, New Age
5. Principles of Organometallic Chemistry, G. E. Coates, M. L. H. Green, P. Powell and K,
6. Principles of Bioinorganic Chemistry, S. J. Lippard and J. M. Berg, University Science Books
7. Bioinorganic Chemistry, I. Bertini, H. B. Gray, S. J. Lippard and J. S. Valentine, University Science Books.
8. Bioorganic, Bioinorganic and Supramolecular Chemistry, P.S. Kalsi, New Age International (P) Limited.

Course Outcomes:

After completion of this course students will able to:

CO-1. Know the basic terms & principles behind coordination compounds & and bioinorganic chemistry.

CO-2. Understand about energy level in an atom, free ions in crystal field, electronic spectra, magnetic properties and metal ligand bonding in transition metal complexes.

CO-3. Focus their aim for future prospects of research in the field of coordination & bioinorganic chemistry.

Semester III Syllabus
Core Course
Course Code: B020902T
Course Title: Pericyclic, Photochemistry & Rearrangement Reactions

Credit 05 **Hour 60**

Course Objectives: After successful completion of the first year of Masters, students coming in III[rd] semester will be provided detailed knowledge about pericyclic, photochemistry & rearrangements and name reactions.

Unit I

Pericyclic Reactions: Basic concepts, classifications, characteristics, conservation of MO symmetry, FO of ethylene, 1-3 butadiene, 1,3,5-hexatriene, allyl system, Woodward, Hoffmann rule for correlation diagram, FMO & PMO approach to study of *Electrocyclic reactions* of linear conjugated 4n, 4n+2, allyl system; *Cycloaddition reaction* of [2+2, [4+2] system, [2+2] addition of ketene, 1, 3-dipolar cycloaddition; *Sigmatropic*: [1,3], [1,5], [3,3], [5, 5] group transfer reaction, suprafacial, antarafacial shift of H, Sigmatropic shift involving carbon moieties, Claisen, Cope rearrangement, Fluxional tautomerism, Aza cope rearrangement, Ene reaction, Chelotropic reaction, Prototropic reaction.

Unit II

Photochemistry: Electronically excited states, spin multiplicity, Jablonski diagram, ISC; *Photochemistry of alkenes*: Geometrical isomerization, cyclisation, dimerization, di-pi methane rearrangement, H abstraction addition, acetylene dimerization, photochemistry of diene, 1,3-butadiene, [2+2] addition leading to cage structure; *Photochemistry of Carbonyls*: Reduction, Norrish I cleavage of acyclic, cyclic, α, β and β, γ unsaturated carbonyl compounds, photochemistry of Norrish II cleavage, Paterno Buchi reaction, intra and inter molecular H abstraction, rearrangement of α, β unsaturated ketones, cyclohexadienones, photoenolization, photocycloaddition of ketones with unsaturated compound, photodimerisation of (enones), α, β unsaturated ketones, rearrangement of enones, dienones photochemistry of p- beinzoquinones.

Unit III

Photochemistry of Aromatics: Ring isomerization, excited state of benzene and its 1,2 & 1, 3 shifts, photo Fries rearrangement (of anilide), cyclisation reactions, Skeletal isomerism, Dewar Prismane isomerization of disubstituted benzene, photo substitution reaction of benzene, photolysis of nitride ester.

Unit IV

Rearrangements: *Photochemical Rearrangements*: Sommlet Hauser, Hofmann-Loffler Freytag, Barton, fries; *Molecular Rearrangements*: General mechanistic considerations, nature of migration, migratory aptitude, memory effects, a detailed study of the following rearrangements- Wagner-Meerwein, Demjanov ring expansion, Dieone-phenol, Benzil-Benzilic acid, Wolf, Lossen,

Beckman, stevens, wittig, Neber, Arndt-Eistert synthesis, Amino ketone, benzidine, Shapiro reaction.

Unit V

Selective Name Reactions: Stark enamine, Chichibebin, Birch reduction, Heck, Suzuki, Mukaiyama, Woodward & Prevost hydroxylation, Peterson synthesis.

Recommended books:

1. Textbook of Pericyclic Reaction, Concept and Application, K.C. Majumdar and P. Biswas, Scientific International Pvt. Ltd.
2. Organic Photochemistry: A visual approach, Jan Kopecky, VCH publishers (1992).
3. Organic Photochemistry, O. Kan, McGraw-Hill Inc., US.
4. Organic Chemistry, J. Clayden, N. Greeves, S. Warren and P. Wothers, Oxford.
5. Fundamentals of Photochemistry, K. K. Rohatagi, New Age International (P) Limited.
6. Photochemistry and Pericyclic Reactions, Jagdamba Singh and Jaya Singh, New Age
7. Principles of Molecular Photochemistry, Nicholas J. Turro, V. Ramamurthy J. C., Viva Books. International (P) Limited.
8. Organic photochemistry, J. Coxon and B. Halten, Cambridge University Press.
9. Essential of Molecular Chemistry, A. Gilbert and J, Baggot, Blackwe II.
10. Molecular photochemistry, N. J. Turro, W. A. Benjamin.
11. Introductory Photochemistry, A. Cox and T. Camp, McGraw-Hill
12. Photochemistry, R. P. Kundalland, A. Gilbert, Thomson Nelson.

Course Outcomes:

CO-1. Students will have knowledge to understand of the various pathway of photochemical, pericyclic, rearrangement and name reactions.
CO-2. Students will able to design the synthetic routes and mechanisms of targeted molecules.
CO-3. Students will focus their aim for future prospects of research in the above area of chemistry.

Semester III Syllabus
Core Course
Course Code: B020903T
Course Title: Electrochemistry

Credit 05 **Hour 60**

Course Objectives: After successful completion of the first year of M.Sc., students, coming in III[rd] semester will be provided detailed knowledge about electrochemistry viz electrokinetic phenomenon, Electrolytic conductance, transference and interface, bio electrochemical cell reaction, polarography & voltammetry, fuel cells & batteries and conductors & semiconductors.

Unit I

Electrokinetic Phenomenon: Electrokinetic Effects, Electrokinetic potential/Zeta potentials, Determination of zeta potential, influence of ions on electrokinetic phenomena, Quantitative treatments of Electro-Osmosis, Electrophoretic and Streaming potential, sedimentation potential, The electrical double layer, Theoretical and quantitative treatment of electrokinetic phenomena, Mobility and Bound hydrogen ion.

Unit II

Electrolytic Conductance, Transference and Interface: Debye Huckel theory of strong electrolyte (DHO eqn.), Debye Falkenhagen effect, Wein effect the ionic association, effect of ionic strength on rate of ionic reactions, activity coefficient, ionic strength, its effect on reaction rate, Debye Huckel theory of mean activity coefficient of strong electrolyte (DHLL); *Electrolyte interface:* Bjerrum theory of ion association in electrolyte solutions, Lippmann equation, determination of surface excess, structure of electrified interface, The Helmholtz –Perin Theory & Guoy–Chapman diffuse charge model of double layer, Stern Modification in the Gouy-Chapman Theory.; *Electrodics*: The equilibrium exchange current density, Butler Volmer Equation, Tafel plot, high field and low field approximation, Electrodeposition and electropolymerization, *Irreversible electrode process*: Overvoltage, corrosion (mechanism, corrosion current, corrosion potential, electrochemical corrosion theory, estimation of corrosion rates prevention methods, polarization resistance, electrodeposition. Polarization Resistance

Unit III

Electrochemical Cell Reactions: Galvanic cells, half reactions and reversible electrodes, single electrode potential, thermodynamics of reversible electrodes and cells, Nernst equation, Standard Electrode potential, Electrochemical series, EMF of Galvanic cells; *Fuel Cells and Batteries:* Fuel cell and its theory, different types of fuel cell, Solid oxide fuel cells (SOFC), Polymer electrolyte fuel cell (PEM), Direct Electrolyte Fuel Cell (DEFC); *Super Capacitors*: Theory, Measurements and importance; *theories of Batteries*: Solid state batteries.

Unit IV

Polarography and Voltammetry: Principle of polarography, variations of the conventional polarographic methods, Pulse Polarography, AC polarography, square wave polarography, Anodic stripping and Cyclic voltammetry, Qualitative and quantitative application of polarography, Determination of stoichiometry and formation constants of complexes, Amperometric titrations and advantages.

Unit V

Conductors and Semiconductors: General principles of semi conductivity and semiconductors, Temperature dependence of electrical resistances, Coherent Length, Piezoelectric effect, Piezoelectric and pyroelectric materials, Fullerenes-Doped conductors, Brief idea of Electrochemistry of molten electrolytes and non-aqueous solvents.

Recommended Books:

1. Modern Electrochemistry, Vol.1&2, J.M. Bockrisand A.K.N Reddy, Plenum.
2. Introduction to electrochemistry, S. Glasston, VanNostrand.
3. Electro-Analytical Chemistry, J. J. Lingane, Willey Inter science.
4. Polarography, D.R. Crow, J. V. Westwood, Methuen and Co.
5. Principle of Polarography, J. Heyrovsky, P. Zuman and L. Kuta
6. Solid state Electrochemistry, Haldil, Academis Press.
7. Electrochemistry of solids, H. Rickett, Springer Book.
8. Ions, Electrodes and Membranes, J. Koryta, Willey and Sons.
9. Electrochemistry, C. W Devis, George Newone, London.
10. Polarography and voltammetry, H.H Bauer & J.E.O Reily.
11. Physical Chemistry, Thomas Engel and Philip Reid, L P E, Pearson Education.
12. Principal of physical chemistry, S.H. Maron and C.F. Prutton, Oxford.
13. Electrode Kinetics, E. Gileadi, VCH Publishers Inc., New York.
14. Electrochemical Methods: Fundamental &applications(2ndEd.), Bard & L. R. Faulkner, John Wiley & Sons, New York
15. Bioelectrochemistry: Fundamentals, Experimental Techniques and Applications, P. N. Bartlett, John Wiley & Sons, Ltd

Course Outcomes:

After completion of this course students will able to:

CO-1. Understand about electrokinetic phenomenon, Electrolytic conductance, transference and interface: bioelectrochemistry, polarography & voltammetry, fuel cells & batteries and conductors & semiconductors.

CO-2. Have knowledge about general principles of semi conductivity, semiconductors, fullerenes-doped conductors, electrochemistry of molten electrolytes and non-aqueous solvents.

CO-3. Develop ideas for further research in the area of electrochemistry.

Semester III Syllabus
Fifth Elective Course
Course Code: B020904T
Course Title: Natural Products

Credit 05 **Hour 60**

Course Objectives: After successful completion of the first year of M.Sc., students, coming in III[rd] semester, if opted this course, will be provided knowledge about various classes natural products in term of their classification, isolations, structural elucidations and synthesis.

Unit I

Alkaloids:, General methods of structure elucidation, classification based on N-heterocyclic ring, Degradation, *Stereochemistry and synthesis*: Nicotine, Quinine, Morphine (Retrosynthesis also), Ephedrine, Reserpine (Retrosynthesis also).

Unit II

Terpenoids and Carotenoids: Classification, General methods of structure elucidation, isoprene rule; *structure, stereochemistry, synthesis*: Camphor (Retrosynthesis also), Abietic acid, Squalene, Citral, α-Terpenol, Menthol, Farnesol, Santonin, β-Carotene, Longifolene (Retrosynthesis also).

Unit III

Steroids: Basic skeleton, Diels' hydrocarbon and stereochemistry, structural determination and synthesis of cholesterol, testosterone, estrone and progesterone.

Unit IV

Prostaglandin: Occurrence, nomenclature, classification, physiological effects and synthesis of PGE_2 and PGF_{2a}.

Unit V

Proteins: Amino acids, polypeptide, structure of protein.
Nucleic acids: General structure of RNA and DNA.
Biosynthesis: Acetate hypothesis, poly-β-keto acid, meta orientation of hydroxyl group in naturally occurring phenols, biogenesis of fatty acids, mevalonic acid from acetyl coenzyme A, bisynthesis of mono, sesqui, di, tri terpenes, shikimic acid pathway for biosynthesis of aromatic ring, general biosynthesis of alkaloids.

Recommended Books:

1. Organic Chemistry, I.L. Finar Vol. I and II, ELBS.
2. Natural Products: Chemistry and Biological, J. Mann. R.S. Davidson, J.B, Hobbs, D.V.

Banthrope and J.B. Harborne, Longman, Essex.
3. Chemistry, Biological and Pharmacological Properties of Medicinal Plants from the Americas,Ed. Kurt Hostettmann. M.P. Gupta and A. Marston, Harwood Academic publishers.
4. Stereoselective Synthesis: A Practical Approach, M. Nogradi, VCH.
5. Rodds Chemistry of Carbon Compounds, Ed. S. Coffey, Elsevier.

Course Outcomes:

After completion of this course students will gain knowledge about:

CO-1. Isolation and structural determination of various classes of natural products.

CO-2. Synthesis and bio synthesis of biomolecules.

CO-3. General structure of nucleic acids and configuration & conformation of carbohydrates.

Semester III Syllabus
Fifth Elective Course
Course Code: B020905T
Course Title: Medicinal Chemistry

Credit 05 **Hour 60**

Course Objectives: After successful completion of the first year of M.Sc., students, coming in III[rd] semester, if opted this course, will be provided detailed understanding about medicinal chemistry viz. classifications, designing, synthesis, uses and mechanisms of actions of the various classes of drug.

Unit I

Drugs Design: Development of new drugs, Structure activity relationship (Structure activity relationship (SAR), Factors affecting bioactivity, Isomerism, bio-isomerism, spatial considerations, Theories of drug activity, Occupancy theory, rate theory, induced fit theory, Quantitative structure activity relationship, History and development of QSAR, concepts of drug receptors (Receptor site theory), elementary treatment of drug receptor interactions, Introduction to combination synthesis in drug discovery, Physico-chemical parameters, Lipophilicity, partition coefficient, electronic ionization constants, steric, Generic medicines.

Unit II

Local Anti-infective Drugs: *Antitubercular drugs and antimalarial drugs*: Introduction and general mode of action, synthesis of sulphonamides, ciprofloxacin, norfloxacin, dapsone, amino salicylic acid, ethionamide, ethambutol, griseofulvin; *4-aminoquinoline derivatives*: chloroquine, santoquine, camoquin, 8- aminoquinoline, primaquine, PAS, Thiosemicarbazones, hydrazides and thiocarbamides.

Unit III

Psychoactive Drugs: CNS depressants, general anesthetics, hypnotics, sedatives, anti-anxiety drugs, benzodiazepines, mental diseases, Antipsychotic drugs. Synthesis of diazepam, alprazolam, trimethadione, barbiturates and glutethimide, Reserpine, Promazine, chlorpromazine, Mepazine.

Unit IV

Synthesis of Sulpha Drugs: Sulphanilamide derivatives, sulphathiazole, sulphadimidine, sulphaguanidine, sulfadiazine; **Anti-HIV drug:** Crixivan and **Cardiovascular Drugs:** synthesis of amyl nitrate, hydrolaxime verapamil, methyldopa and diazoxide propanol.

Unit V

Structure Based Drugs Classification: *Substituted benzene ring*: Chloramphenicol, salmeterol, Tolazamide, dichlophenac, tiapamil, intriptyline; *Five membered heterocycles*: Tolmetin, Spiralpril, oxaprozine, sulconazole, nizatidine, imolamine, isobuzole; *Six membered heterocycles*: Warfarin,

quinine, norfloxacin, ciprofloxacin, methylclothiazide, citrine, terfenadine; *Seven membered heterocyclic ring fused to benzene*: Chlordiazepoxide, diazepam, diltiazem; *Heterocycles fused to two benzene rings*: Quinacrine, tacrine; *Five membered heterocycles fused to six membered rings*: Acyclovir, methotrexate.

Recommended Books:

1. Medicinal Chemistry, D. Sriram, P. Yogeeswari, Pearson Education.
2. Medicinal Chemistry, Ashutosh Kar, New Age International (P) Limited.
3. An Introduction to Medicinal Chemistry, Graham L. Patrick, Oxford University Press.
4. Textbook of Medicinal Chemistry, V. Alagarsamy, Elsevier Health Sciences.
5. The Practice of Medicinal Chemistry, Camille G. Wermuth, Elsevier Health Sciences.
6. Drug-like Properties: Concepts, Structure Design and Methods: From ADME to Toxicity Optimization, Edward H Kerns, Li Di, Elsevier Health Sciences.
7. Ram V. J.; Sethi, A.; Nath, M.; Pratap, R.; (2019), The Chemistry of Heterocycles (Nomenclature and Chemistry of three to five membered Heterocycles), Elsevier publication.
8. Ram V. J.; Sethi, A.; Nath, M.; Pratap, R.; (2019), The Chemistry of Heterocycles (Chemistry of six to eight membered N, O, S, P and Se heterocycles), Elsevier publication
9. A. Burger; Medicinal Chemistry.

Course Outcomes:

CO-1. Students will train about the synthesis, applications of the various classes of drugs.
CO-2. Students will have understanding of retrosynthetic analysis thus they can prepare the desired molecules.
CO-3. Students will motivate to develop ideas for further research into the medicinal chemistry.

Semester III Syllabus
Sixth Elective Course
Course Code: B020906P
Course Title: Chemistry Laboratory Course-IIIA

Credits-5 **10h/Week**

Course Objectives: The objective of this course to provide advanced insight about the inorganic preparations in aqueous and organic medium, colorimetry, spectrophotometry, flame photometry, qualitative analysis of ternary organic mixtures and three steps organic synthesis, interpretation of the spectral data of organic & inorganic compounds and experiments related to physical chemistry viz. potentiometry and conductometry.

INORGANIC CHEMISTRY

Inorganic preparation in aqueous and organic medium:

i. Preparation of $K_3[Fe(C_2O_4)_3].3H_2O$
ii. Preparation and separation of cis and trans –$[Co(en)Cl_2]$
iii. Preparation of $CuCl_2$ DMSO and Copper glycine complex.
iv. Preparation of Ph_3P and its complexes.
v. Preparation of ferrocene.
vi. Preparation of $Mn(gly)_3$

Spectroscopy: Record the spectra (UV, IR & Mass) and analyze the spectral data of the synthesized inorganic complex compounds.

Colorimetry and Spectrophotometry:

i. Estimation of the metals in solution V, Mo, and Fe Colorimetry /Spectrophotometry.
ii. Colorimetric and Spectrophotometric analysis: Determination of iron, copper, ammonium, phosphate, fluoride and nitrite ions.
iii. To verify Lambert's Beer Law colorimetrically/ spectrophotometrically of inorganic compounds.

Flame Photometry:
i. Estimation of magnesium and calcium in tap water
ii. Estimation of calcium in calcium salt solution

ORGANIC CHEMISTRY

Qualitative Analysis: Separation purification and identification of components of ternary organic mixtures (all liquids, two liquids & one solid, all solids). Each component should not contain more than two functional groups. The student should check the purity by TLC, systematic analysis of each component leading to their final identification laying emphasis on solubility, element detection, melting point, boiling point determination, ignition test, unsaturation test, functional group test and preparation of suitable derivative.

Organic Synthesis: Three step synthesis of organic compounds.
Spectroscopy: Record the spectra (UV, IR, ^{1}H NMR, ^{13}C-NMR and Mass) and analyze the spectral data of synthesized organic compounds.

PHYSICAL CHEMISTRY

Potentiometry:

i. Determination of the solubility of a sparingly soluble salt in water by EMF method.
ii. Titration of ferrous ammonium sulphate against $K_2Cr_2O_7$ (or $KMnO_4$) potentiometrically and determine the formal redox potential of Fe^{2+}-Fe^{3+} system.
iii. Find out the normality of the given HCl solution by titrating it potentiometrically with N/15 NaOH solution.

Conductometry:

i. Determination of the strength of strong acid conductometrically by using strong alkali solution.
ii. Determination of the strength of weak acid conductometrically by using strong alkali solution.
iii. Determination of cell constant of the conductivity meter with the help of KCl solution.

System of Marking:-
Time: 12h
Inorganic: 33
Organic: 34
Physical: 33

Recommended Book:

1. Experimental Inorganic Chemistry by Mounir A, Malati, Horwood series in Chemical Science (Horwood publishing Chichester) 1999.
2. Inorganic Experiments, J. Derexwoolings VCH
3. Practical Inorganic Chemistry, G. Marrand, B.W. Rockett, Van Nostrand.
4. Synthesis and characterization of Inorganic compounds, W.L. Jolly, Prentice Hall
5. The systematic Identification of Organic Compounds, R.L. Shringer and D.Y. Curlin.
6. Handbook of Organic Analysis Qualitative and Quantitative, H. Clark, Adward Ar.
7. Vogel's Textbook of Practical Organic Chemistry, A.R. Tatchell, John Wiley.
8. Systematic Qualitative Organic Analysis, H. Middeton, AdwardArnoid.
9. Practical Physical Chemistry, A.M. James and F.E. Prichand, Longman.
10. Findley's Practical Physical Chemistry revised, B.P. Levitt, Longman.
11. Experimental Physical Chemistry, R.C. Das and Bebera, Tata Mc Grawhill.
12. Senior Practical Physical Chemistry, B.D. Khosla and V.S. Barg (R. Chand and Co., Delhi)
13. Experimental Physical Chemistry by D.P. Shoemaker Mc Grawhill, 7th Edition 2003.
14. Systematic Experiment in chemistry, Arun Sethi, New Age International (P) Limited.
15. Practical Physical Chemistry, B. Vishwanathan and P. S. Raghwan, Viva Books.

16. Experimental Physical Chemistry, V. D. Athawale, Parul Mathur, New Age International (P) Limited.
17. Experiments in Physical chemistry, J.C. Ghosh, Bharati Bhavan.
18. Advanced Practical Physical Chemistry, J B Yadav.

Course outcomes:

After completion of this course students will acquire the knowledge of:

CO-1. Synthesis of inorganic compounds in organic and aqueous medium.

CO-2. Calorimetry, spectrophotometry and flame photometry estimation of metals in solution/ water,

CO-3. Qualitative analysis of ternary organic mixtures and three steps organic preparations.

CO-4. Interpretation of the spectral data of inorganic and organic compounds.

CO-5. Various theories and principles behind the experiments of physical chemistry viz. potentiometry and conductometry.

Semester III Syllabus
Sixth Elective Course
Course Code: B020907P
Course Title: Chemistry Laboratory Course-IIIB

Credit 05 **10h/Week**

Course Objectives: The objective of this course to provide the detailed knowledge about the quantitative estimations in the area of inorganic chemistry, quantitative estimation in the field of organic chemistry, chromatographic separations of ions by using suitable ion exchangers, extraction of compounds from natural sources, and experiments related to physical chemistry viz. p^H-metry and spectrophotometry.

INORGANIC CHEMISTRY

Gravimetric Estimations: Gravimetric estimations of complex mixtures involving two or three constituents, Analysis of alloys and minerals.

Volumetric Estimations:

i. EDTA titrations - Determination of Zn, Ca, Mg and Fe.
ii. $KBrO_3$ and KIO_3 titrations –Determination of As_2O_3 and $[Fe(CN)_6]^{4-}$.
iii. Chloramine T – titrations - Determination of NO_2 in a sample.
iv. Ceric Ammonium Sulphate titrations - Determination of Fe and organic acids.

Chromatography:

(i) Separation of Cl^- and Br^- by suitable ion exchangers
(ii) Separation of Co^{2+} and Ni^{2+} by suitable ion exchangers
(iii) Separation of Zn^{2+} and Mn^{2+} by suitable ion exchangers

ORGANIC CHEMISTRY

Quantitative Analysis:

i. Determination of percentage or number of hydroxyl group in an organic compound by acetylation method.
ii. Determination of percentage or number of amino groups in an organic compound by acetylation method.
iii. Determination of percentage or number of methoxy group in the organic compound by Ziesel's method.
iv. Estimation of amines/phenols using bromate bromide solution/or acetylation method.
v. Determination of iodine and saponification value of an oil sample.
vi. Estimation of Nitrogen in the organic compound by Kjeldahl's method.

vii. Estimation of sulphur in the organic compound by Messenger's method.

Extraction: Isolation of biomolecules from natural sources.

PHYSICAL CHEMISTRY

p^H-metry:

i. Determination of the strength of strong acid with strong base by using p^H meter.
ii. Determination of the strength of weak acid with strong base by using p^H meter.
iii. Verification of Henderson's equation by using p^H meter.

Spectrophotometry:

i. Estimation of the following metals in solution Cr and Ni.
ii. Determination of stability constant of a metal ligand complex by spectrophotometric method.
iii. Investigation of reaction between potassium per-sulphate and potassium iodide by spectrophotometer method.
iv. To verify Lambert's Beer Law spectrophotometrically.

System of Marking:-
Time: 12h
Inorganic: 33
Organic: 34
Physical: 33

Recommended Books:

1. Vogels Text book of Quantitative Analysis revised, J. Bessett, R.C. Denney, G.H. Jellery and J. Mendhan ELBS
2. Experimental Inorganic Chemistry by Mounir A, Malati, Horwood series in Chemical Science (Horwood publishing Chichester) 1999.
3. Inorganic Experiments, J. Derexwoolings VCH
4. Practical Inorganic Chemistry, G. Marrand, B.W. Rockett, Van Nostrand.
5. Handbook of Organic Analysis Qualitative and Quantitative, H. Clark, Adward Ar
6. Systematic Experiment in chemistry, ArunSethi, New Age International (P) Limited.
7. Practical Organic Chemistry, Mann and Saunders.
8. Vogel's Textbook of Practical Organic Chemistry, A.R. Tatchell, John Wiley.
9. Practical Physical Chemistry, A.M. James and F.E. Prichand, Longman.
10. Findley's Practical Physical Chemistry revised, B.P. Levitt, Longman.
11. Experimental Physical Chemistry, R.C. Das and Bebera, Tata Mc Grawhill.
12. Senior Practical Physical Chemistry, B.D. Khosla and V.S. Barg (R. Chand and Co., Delhi)
13. Experimental Physical Chemistry by D.P. Shoemaker Mc Grawhill, 7th Edition 2003.
14. Experiments in Chemistry, D.V. Jahagirdar, Himalaya Publishing House.
15. Practical Physical Chemistry, B. Vishwanathan and P.S. Raghwan, Viva Books.

16. General Chemistry Experiments, Anil J Elias, University Press (2002)
17. Experimental Physical Chemistry, V.D. Athawale, ParulMathur, New Age International (P) Limited.
18. Experiments in Physical chemistry, J.C. Ghosh, BharatiBhavan.
19. Advanced Practical Physical Chemistry, JB Yadav.
20. Practical Organic Chemistry, Mann and Saunders.

Course outcomes:

After completion of this course students will acquire the knowledge of:

CO-1. Quantitative estimation of the complex mixtures containing two or three constituents and analysis of alloys and minerals.
CO-2. Chromatographic separations of ions by using suitable ion exchangers.
CO-3. Estimations of functional groups and element present in the organic compounds.
CO-4. Extraction of biomolecules from natural sources.
CO-5. Various theories and principles behind the experiment of Ph-metry and spectrophotometry.

Semester IV Syllabus
Core Course
Course Code: B021001T
Course Title: Organotransition Metal Chemistry

Credit 05 **Hour 60**

Course Objectives: After successful completion of the III[rd] semester of M.Sc., students coming in IV[th] semester will be provided knowledge about the compounds having metal carbon bond & metal hydrogen bond, organotransition metal catalyst, macrocyclic complexes and fluxional organometallic compounds.

Unit I

Metal Carbonyls: Preparation, Structure and reactions/ properties of mononuclear and polynuclear metal carbonyls, Nature of M-CO bonding. Vibrational spectra of metal carbonyl, Principal reaction types, Verities of CO bridging; *Metal nitrosyls:* bonding, structure, Metal carbonyl-metal nitrosyl complex: Carbonyl metal hydride. vibrational spectra of metal carbonyls for structural elucidation.

Unit II

Organometallics: Nomenclature, general characteristics, Major types of transition metal to carbon bonds, Preparation stability and important reaction of transition metal alkyl and aryls; *Inorganic π-Acid ligands:* O, N, tertiary phosphine and arsines as ligands ; *Complexes of σ-donor ligands:* General methods of preparation, properties, nature of bonding and structural features of Transition metal alkenyls, alkynyls and carbines, carbines & Pi Complexes of unsaturated molecules: alkenes, alkynes, allyl, dienes, dienyl cyclopentadienyl, thenyl (arenes) complexes, Important reactions related to nucleophilic and electrophilic attack on ligands, reactions, with special reference to organic synthesis; *Transition Metal compounds with M-H bond:* Metal hydrides (Classical, non-classical), synthesis and important reactions; *Metal alkoxides:* Preparation, Properties, Structure, Industrial application.

Unit III

Organometallic Catalyst: General ideal of important catalytic steps, ligands coordination, and dissociation, and elimination, nucleophilic attack on coordinated ligands & coordinated molecular oxygen, Template synthesis, Oxidative addition, Reductive elimination and migration (insertion) reactions; *Homogeneous Catalysis:* Hydrogenation of alkenes using Wilkinson's catalyst, Hydroformylation of alkenes using Co and Rh catalysts, Carbonylation of methanol to acetic acid (Monsanto process), Oxidation of alkenes (Wacker process).

Unit IV

Metal Clusters & Micro–Macrocyclic Complex: M-M multiple bonds containing binuclear, trinuclear, tetranuclear and octahedral clusters, synthesis and bonding in clusters, metal carbonyl halides, Chalcogenide clusters; Types of macrocyclic ligands, design and synthesis by coordination template effect, di and polynuclear macrocyclic complexes, Application of macrocyclic complexes.

Unit V

Fluxional Organometallic Compounds: Fluxionality and dynamic equilibria in compounds such as η^2-olefine, α allyl and dienyl complexes; *Organometallic Compounds of Lanthanides and Actinides*: Methods of preparation, properties and structural features.

Recommended Books:

1. Advanced Inorganic Chemistry, F. A. Cotton and G. Wilkinson, John Wiley
2. Inorganic Chemistry, J. E. Huheey, Ellen A. Keiter, Richard L. Keiter, Addison Wesley Longman (Singapore) Pvt. Ltd.
3. Chemistry of the Elements, N. N. Greenwood and A. Earnshow, Pergamon.
4. Organometallic Chemistry: A Unified Approach, R. C. Mehrotra and A. K. Singh, New Age
5. Principles of Organometallic Chemistry, G. E. Coates, M. L. H. Green, P. Powell and K. Wade, Chapman and Hall, London.
6. Comprehensive Organometallic Chemistry, Ed. E.W. Abel, Abel, F.G.A. Stone and G. Wilkinson, Pergamon.

Course Outcomes:

After completion of this course students will be able to:

CO-1: Understand detailed about metal carbonyls & organotransition metal complexes having M-C & M-H bonds.

CO-2: Acquire understanding of organotransition metal complex having π acids ligands.

CO-3. Have understanding of organometallic catalysts, macrocyclic complex and fluxional organometallic compounds.

CO-4. Develop idea for further research in the field of organotransition metal chemistry.

Semester IV Syllabus
Core Course
Course Code: B021002T
Course Title: Organic Synthesis

Credit 05 **Hour 60**

Course Objectives: After successful completion of the III[rd] semester of M.Sc., students, coming in IV[th] semester will be provided detailed understanding about organic synthesis viz. applications of Oxidation reduction process, reagents in organic synthesis, asymmetric synthesis, retrosynthetic analysis and synthesis of some complex molecules and green chemistry.

Unit I

Oxidation & Reduction: Different oxidative processes, Hydrocarbon alkenes, aromatic ring, saturated C-H groups (activated and inactivated) Alcohols, diols, aldehydes, ketones, carboxylic acids, amines, hydrazine, sulphide; *Reduction*: Different reductive process, Hydrocarbon-alkanes, alkenes, alkynes, aromatic rings, carbonyls-aldehyde, ketones, acids, acid derivatives, epoxides, hydrogenolysis.

Unit II

Reagents: $LiAlH_4$, $NaBH_4$, $SnBu_3H$, $RhCl(PPh)_3$, $IC_6H_5(OAc)_2$, SeO_2, RuO_4, OsO_4, RCO_3H, HIO_4, $Pb(OAc)_4$, CH_2N_2, NBS, R_2CuLi, LDA, DCC, 1,3-dithiane (reactivity umpolung), Me_3SI, Baker's yeast, organophosphorus compounds, ylides (S, N, P), Phase transfer catalyst, quaternary ammonium and phosphonium salts, crown ethers, Merrifield resins, DDQ, Jone's reagent, $TI(NO_3)_3$, DIBAL, B_2H_6, di-isoamylborane, 9-BBN.

Unit III

Asymmetric Synthesis: Stereospecific, stereoselective synthesis, Enzymatic and catalytic nexus, Enantioselective synthesis with chiral non racemic and catalysts, hydroboration with chiral boranes ($I_{PC}BH_2$). $(A_{PC})_{21}BH$, carbonyl group reductions and chiral complex hydride (BINAL-H). Chiral oxazaberlidines, Diastereoselective synthesis, Asymmetric synthesis involving chiral, auxiliary chiral reagent and chiral catalysis, methods of resolution, enantiomeric excess i.e., quasi racemate and optical purity.

Unit IV

Retrosynthetic Analysis: Synthon, synthetic equivalent, one group C-X and two group C-X disconnection, Disconnection (C-C, C-S, C-O) bonds, FGI, Chemoselectivity, Cyclisation reactions, synthetic strategy for formation of C-C, C-N, C-X bonds. Reversal of polarity, Amine synthesis, multistep synthesis; *Protection*: Principles, deprotection of alcohols, thiols, 1,2 and 1,3-diols, amines, carbonyls and carboxyl groups in organic synthesis.

Unit V

Synthesis of Some Complex Molecules and Green Chemistry: Application of the above in the synthesis of following compounds: Camphor, Longifolene, Cortisone, Reserpine, Vitamin D, Juvabione, Aphidicolin and Fredericamysin A; *Green Chemistry:* Basic Principle, Microwave assisted organic synthesis, Combinatorial chemistry.

Recommended Books:

1. H.O. House, Synthetic Organic Chemistry, Benjamin-Cummings Publishing Co.
2. Organic Chemistry J. Clayden, N. Greeves, S. Warren and P. Wothers, Oxford Press.
3. Organic Synthesis, Pragati Edition, Jagdamba Singh and L.D.S Yadav.
4. Some modern methods of organic synthesis, W. Carruthers, Cambridge University Press.
5. Organic Reactions and Their Mechanisms, P. S. Kalsi, New Age Science.
6. Workbook for Organic Synthesis, Stuart Warren, John Wiley & Sons.
7. Graham Solomon, T.W., Fryhle, C.B. & Dnyder, S.A. Organic Chemistry, John Wiley & Sons.
8. Ram V. J.; Sethi, A.; Nath, M.; Pratap, R.; (2019), The Chemistry of Heterocycles (Nomenclature and Chemistry of three to five membered Heterocycles), Elsevier publication.
9. Ram V. J.; Sethi, A.; Nath, M.; Pratap, R.; (2019), The Chemistry of Heterocycles (Chemistry of six to eight membered N, O, S, P and Se heterocycles), Elsevier publication.

Course Outcomes:

CO-1. Students will train about the handling and applications of reagents in organic synthesis.
CO-2. Students will able to understand designing and synthesis of targeted molecules by applying retrosynthetic approach.
CO-3. Students will train to develop ideas for further research in the field of synthetic organic chemistry.

Semester IV Syllabus
Seventh Elective Course
Course Code: B021003T
Course Title: Analytical Chemistry

Credit 05 **Hour 60**

Course Objective: After successful completion of the III[rd] semester of M.Sc., students, coming in IV[th] semester, if opted this course, students have to study the various theories, concepts and principles behind the instrumental techniques which can be used to solve various intricate problems in science.

Unit I

Data Analysis: Systematic and random errors, Accuracy and Precision and ways of its expression, Fitting data to a straight line, Normal error curve and its equation, Propagation of errors, Standard tests, Test of significance, F-test, Student t test, Q-test, Chi- test, Correlation test, distribution normalcy test, confidence limit of mean, comparison of two standard value, comparison of standard deviation and average deviation, comparison of mean with true value, significant figures, Rounding of figures, Regression analysis (least square method for linear plots), Rejection of observations, Statistics of sampling and detection limit evaluation.

Unit II

Electroanalytical Techniques: Conductometry, Polarography (Limiting current density, Dropping mercury electrode, Ilkovic equation, Half wave potential), Voltammetry, Cyclic voltammetry, Anodic stripping voltammetry, Potentiometry, Amperometry, Ion selective electrodes, Coulometry.

Unit III

Thermoanalytical Techniques: Apparatus, factors affecting TG, Interpretation of TG curves of $CaC_2O_4.H_2O$ and $MgC_2O_4.2H_2O$. Apparatus, factors affecting DTA and DSC curves with special reference to heating rate, particle size and packing, measurement of heat of transition, heat of reaction and heat of dehydration of salts and metal hydrates.

Unit IV

Spectroanalytical Techniques: Elementary idea of nephleometry and Turbidimetry.

Unit V

Separation Techniques: *Distribution law*: Principles and application of solvent extraction, *Chromatography*: Column, ion exchange and size exclusion chromatography. GLC, GSC, HPLC, electrophoresis.

Recommended Books:

1. Lloyd R. Snyder LC Resoures, Inc walnut Greek, California
2. Colin F. Poole, Department of Chemistry, Wayne State University Detroit MI 48202, USA 2003, Elsevier.
3. J. D. Seader, and Ernest J. Henley, Separation Process Principles, Wiley, 2nd edition (2013).
4. Fundamentals of analytical Chemistry, D. A. Skoog, D. M. West and F. J. Holler.
5. Analytical Chemistry, Theory practice, U.N. Das, Sultan Chand and Sons, New Delhi.

Course Outcomes:

After completion of this course:

CO -1: Student will be able to know the basic concept involved in the analytical chemistry.

CO-2: Students will gain the understanding of data analysis, electroanalytical, thermo analytical, Spectro analytical and spectral techniques which is frequently involved in various research areas.

Semester IV Syllabus
Seventh Elective Course
Course Code: B021004T
Course Title: Polymer Chemistry

Credit 05 **Hour 60**

Course Objectives: After successful completion of the III[rd] semester of M.Sc., students, coming in IV[th] semester, if opted this course, will be provided knowledge about the principles, synthesis, structures, chemical and physical properties of polymers.

Unit I

Basic Concept & Polymer Characterization: General definition, Monomers, repeating units, degree of polymerization, Types and Classification of polymers, Concept of average molecular weights in polymers: (Number average, Weight average, Viscosity average and Sedimentation average molecular weights), Concepts of Mono-dispersity, poly-dispersity, Significance of Molecular Weight, Distribution Curves of polymers, The practical significance of molecular weight; *Measurement of molecular weights*: End-group, viscosity, light scattering, osmotic and ultracentrifugation methods.

Unit II

Kinetics & Mechanism of Polymerization: Polymerization conditions and polymer reactions, Polymerization in homogeneous and heterogeneous systems, Kinetics and mechanism of condensation, Addition (Radical chain and Ionic chain), Coordination and Copolymerization.

Unit III

Rheology & Degradation: *Rheology:* Viscous flow (Newtonian and Non- Newtonian fluids), Rubber elasticity (thermodynamic and entropy, elasticity), Visco-elasticity, The glassy state and glass transition temperature; *Degradation:* Types of degradation: Random degradation and Chain depolymerisation, A general idea of thermal, mechanical and oxidative degradation, Antioxidants and stabilizers.

Unit IV

Polymer Processing: Plastics, elastomers and fibers, Compounding, *Processing techniques*: Calendaring, die casting, rotational casting, film casting, Injection moulding, blow moulding, extrusion moulding, thermoforming & thermofoaming, reinforcing and fiber spinning.

Unit V

Properties of Commercial Polymers: Polyethylene, polyvinyl chloride, polyamides, polyesters, phenolic resins, epoxy resins and silicone polymers; *Functional polymers*: Fire retarding polymers

and electrically conducting polymers; *Biomedical polymers*: contact lens, dental polymers, artificial heart, kidney, skin and blood cells; *Light emitting polymers:* Elementary idea of light emitting polymers.

Recommended Books:

1. Textbooks of Polymer science, F.W. Billmeyer, Jr. Wiley.
2. Polymer Science, U.R. Gowariker, N.V. Vishwanathan and J. Sreedhar, Wiley-Estern.
3. Functional Monomers and Polymers, K. Takemoto, Y.Inaki and R.M. Ottanbrite.
4. Contemporary Polymer Chemistry, H. R. Allcock and F.W. Lambe, Prentice hall.
5. Physics and Chemistry of Polymers, J.M.G. Cowie, Blackie Academic and Professional

Course Outcomes:

After completion of this course students will be able to:

CO-1. Have knowledge about the basic concepts and principles of polymers.
CO-2. Understand about polymer chemistry- characterization, polymerization, and kinetics of polymerizations, degradation, rheology and processing.
CO-3. Develop ideas for further research in the area of polymer chemistry.

Semester IV Syllabus
Research Project/Dissertation
Course Code: B021001P
Course Title: Practical Based Major Research Project/Dissertation

Credits 10

Course Objectives: The objective of this course is to develop interest of students towards major research project/ dissertation and to elevate their understanding towards experimental aspect of some targeted fields of Chemistry. Analytical ability of the students is to be developed and to train the students to work in any research group by motivating them to execute research in the area of their interest in chemical sciences.

Course Outcomes:

CO-1. After completing this major research project/dissertation, students will learn to work independently.
CO-2. Students will be able to plan and strategize a scientific problem, and implement it within a reasonable time frame.
CO-3. Students will be able to know the library search and to interpret the spectral data independently.
CO-4. Students will be able to critically examine research articles, and will improve their scientific writing as well as communication skills.
CO-5. Students will be able to present their finding by using OHP/PPT.

For major research project work/dissertation, the area of the work would be decided by the advisor/mentor/HOD. On completion of the major research project work/dissertation, students have to submit the work in the form of a dissertation followed by oral presentation in the presence of faculty members and finally it will be evaluated by internal and external examiners followed by conducting viva voce examination.